Chemical Changes Homework Week 6

In your jotter write the DATE in the margin.
Now add a TITLE for your homework.
Your title should be the name of the Homework: Chemical Changes Homework Week 6

Question 1.

Which element is missing from the word equation?

[ ] + oxygen ⟶ hydrogen oxide

A carbon
B copper
C hydrogen
D oxygen

Question 2.

Sulfur	
Boiling point	445 °C
Melting point	113 °C

When sulfur cool from 120°C to 100°C
it changes from

A liquid to gas
B gas to liquid
C liquid to solid
D solid to liquid.

Question 3.

One of the products of the reaction between sulphuric acid and copper oxide is copper sulphate. The other product is water.

The word equation for the reaction is

A copper sulphate + copper oxide → sulphuric acid + water
B copper sulphate + water → sulphuric acid + copper oxide
C sulphuric acid + copper sulphate → copper oxide + water
D sulphuric acid + copper oxide → copper sulphate + water

Question 4.

A pupil reacted 3g of calcium carbonate with 50 cm^3 of dilute hydrochloric acid at room temperature.

Suggest 2 ways in which the pupil could increase the rate of reaction.

Question 5.

Below are some example of everyday chemical reactions. Use your knowledge of the factors that affect reaction rate to explain the following.

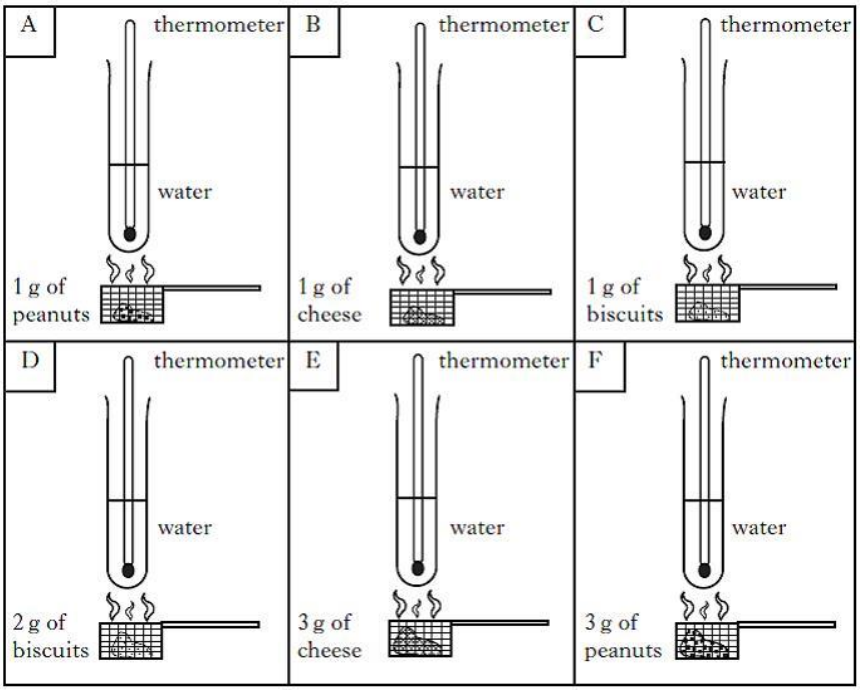
(a) Why do powders that neutralize acid indigestion in your stomach work faster than tablets?

(b) Explain why food is kept fresher for longer if stored in a refrigerator.

(c) Explain why car exhausts rust faster than the underside of cars.

Question 6.

Sajeed investigated the energy content of foods. He set up six experiments. In each experiment, he burned the food and timed how long it took for the water temperature to rise by 25 ºC. Which two experiments should Sajeed compare to find out whether peanuts or biscuits have more energy?



A A & C
B A & D
C F & C
D F & D

A marking guide for this Homework is available (password required).