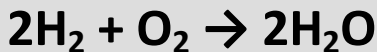


Name _____ Date Due _____

Quantities from Equations

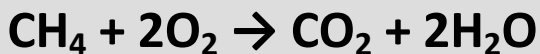


1. From the equation above...

a. Use $n = m \div \text{GFM}$ to work out the number of moles in 30 g of hydrogen gas.

b. How do we know that the number of moles of water produced from burning 30 g of hydrogen gas will be the same as the answer to (a)?

c. Use $m = n \times \text{GFM}$ to work out the mass of water produced from burning 30 g of hydrogen gas.



2. From the equation above...

a. Use $n = m \div \text{GFM}$ to work out the number of moles in 100 g of methane (CH_4) gas.

b. How do we know that the number of moles of water produced from burning 100 g of methane gas will be the twice the answer to (a)?

c. Use $m = n \times \text{GFM}$ to work out the mass of water produced from burning 100 g of methane gas.