

Name _____ Date Due _____

Moles to Mass

1. Use $m = n \times \text{GFM}$ to work out the mass of...

1 mole of H_2

3 moles of NH_3

0.34 moles of CH_4

1.8 moles of $\text{C}_2\text{H}_6\text{O}$

104 moles of $\text{C}_6\text{H}_9\text{N}_3\text{O}_2$

2. Use $n = m \div \text{GFM}$ to work out the number of moles in...

2 g of H_2

140 g of NH_3

0.5 g of CH_4

1.4 g of $\text{C}_2\text{H}_6\text{O}$

40 g of $\text{C}_6\text{H}_9\text{N}_3\text{O}_2$