Summary - "Chemicals"

Signs of a Chemical Reaction

The three main signs that a chemical reaction has happened are a colour change, solid produced (goes cloudy) or a gas is given off (fizzing/bubbles).

Chemical Reactions and Energy

A chemical reaction often turns **chemical energy** into **light, sound** and **heat** energy.

Elements

The **elements** are the **simplest "building blocks"** that make up us, the earth and the universe. They are listed in the **Periodic Table**.

Metals and Non-Metals

You can mark the **metals** and the **non metals** by drawing a **staircase** down the periodic table (see over). **Metals conduct** electricity. Most **non-metals do not conduct** electricity **apart from carbon** which does.

Groups (columns) of elements

The elements are arranged in "families" or "groups" which are the columns in the Periodic Table. The have numbers from 1-8/0. Some of the families also have names.

Alkali Metals

The **alkali metals** are **group 1** in the Periodic Table. The **react violently with water** producing **hydrogen gas** and the **hydroxide** of the alkali metal.

e.g. lithium + water \rightarrow lithium hydroxide

Noble Gases

The **noble gases** are group **8/0** of the Periodic Table. Most of the time, they **do not react with anything** which is why they are called "noble". They are often used in lights

Halogens

The **halogens** are **group 7** of the Periodic Table. They are quite reactive.

Compounds and Mixtures

A compound is more than one element joined together. In a mixture, the elements are not joined together. Compounds with only two elements in them get their names from the name of the two elements with "ide" at the end. If there is a metal, it is written first. If a compound name ends in "ate", then it contains oxygen.

- e.g. iron + oxygen \rightarrow iron oxide
- e.g. sulphur + chlorine \rightarrow sulphur chloride
- e.g. iron + carbon + oxygen \rightarrow iron carbonate

Writing down Chemical Reactions

A chemical reaction can be **written** with the **reactants** (things you start with) on the **left** and the **products** (things you end up with) on the **right** with an **arrow** pointing between them.

e.g. hydrogen + oxygen \rightarrow hydrogen oxide

The Periodic Table

→ Non-metals

Metals ◆

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Group 8/0 (Noble Gases) 2 He	10 Neon	18	¥.	Argon 36	ᅔ	Krypton	24	, ×	98	R	Radon	118	Uuo
Group 7 (halogens)	9 F	17	ָּס	35	Br	Bromine	53	- :	lodine 85	Ąţ	Astatine	117	Uus Ununseptium
Group 6	8 Oxygen	16	S	34 34	Se	Selenium	25	Te	Tellurium 84	Po	Polonium	116	÷
Group 5	7 N Nitrogen	15	۵ :	Phosphorus 33	As	Arsenic	51	qs :	Antimony 83	Bi	Bismuth	115	Uup Ununpentium
Group 4	6 C Carbon	14	Si	32	Ge	Germanium	20	Sn	1 8 8 8 8 8 8 8 8 8 8 8 8 8 8 8 8 8 8 8	Pb	Lead	114	FI Flerovium
Group 3	5 B Boron	13	₹	31	Ga	Gallium	49	ِ ع	Indium 81	F	Thallium	113	Uut Ununtrium
				30	Zn	Zinc	48	ප <u>;</u>	Cadmium 80	Hg	Mercury	112	Copernicium
				29	J	Copper	47	Ag	Silver 79	Αn	Gold	111	Roentgenium
-	I			28	Ē	Nickel	46	Pd	Palladium 78	Pt	Platin um	110	DS Darmstadtium
1 T				27	ಽ	Cobalt	45	윤	Rhodium 77	۲	Iridium	109	Meitnerium
				26	Fe	Iron	44	Ru	Kuthenium 76	Os	Osmium	108	HS
S				25	Σ	Manganese	43	<u>ٔ</u> ۲	Technitium 75	Re	Rhenium	107	Bh Bohrium
quid	Gases			24	ပ်	Chromium	42	β	Molybdenum 74	>	Tungsten	106	Sg Seaborgium
ت	<u>ٿ</u>	5	1	23	>	Vanadium	41	g	Niobium 73	Та	Tantalum	105	Db
				22	F	Titanium	40	Zr	Zirconium 72	Ŧ	Hafnium	104	KT Rutherfordiu
				21	Sc	Scandium	39	>	Yttrium 71	3	Lutetium	103	Lr Lawrencium
Group 2	4 Beryllium	12	Mg	Magnesium 20	ca	Calcium	38	ا	strontium 56	Ва	Barium	88	Ra Radium
Group 1 (alkali metals)	3 Li	11	Na ,	30dium 19	¥	Potassium	37	g. ¦	Rubidium 55	ర	Caesium	87	Fr Francium

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	n Thulium	101	E P	m Mendeleviu
68 Fr	ш	100		n Fermium
	Holmium	66	Es	Einsteinium
99 2	Dysprosium	86	ರ	Californium
65 Tb	Terbium	6	쓢	Berkelium
64 Gd	Gadolinium	96	ဦ	Curium
63 Fu	Europium	92	Am	Americium
62 Sm	Samarium	94	Pa	Plutonium
61 Pm		93	N	Neptunium
09 PN	Neodymium	95	_	Uranium
s9 Pr	Praseodymiu	91	Pa	Protactinium
28 Ce	Cerium	06	ᄕ	Thorium
57 A	Lanthanum	68	Ac	Actinium